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problem, the first

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each reactant present. In this case, we are given the mass of $K_2Cr_2O_7$ in 1 mL of solution, which we can use to calculate the number of moles of $K_2Cr_2O_7$...

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a. If you use 2.3 moles of Al, how many moles of Al_2O_3 can you make? b. If you

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want 3.9 moles of Al_2O_3 , how many moles of O_2 are needed? 2. In the presence of sulfuric acid, metallic iron forms iron(III) sulfate: $2\text{Fe} + 3\text{H}_2\text{SO}_4 \rightarrow \text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2$

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